

A



Surname _____

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I declare this is my own work.

GCSE

COMBINED SCIENCE: TRILOGY

Foundation Tier

Chemistry Paper 1F

F

8464/C/1F

Thursday 14 May 2020

Morning

Time allowed: 1 hour 15 minutes

At the top of the page, write your surname and other names, your centre number, your candidate number and add your signature.

[Turn over]



JUN 20 8464 C 1 F 0 1

For this paper you must have:

- **a ruler**
- **a scientific calculator**
- **the periodic table (enclosed).**

INSTRUCTIONS

- **Use black ink or black ball-point pen.**
- **Pencil should only be used for drawing.**
- **Answer ALL questions in the spaces provided.**
- **If you need extra space for your answer(s), use the lined pages at the end of this book. Write the question number against your answer(s).**
- **Do all rough work in this book. Cross through any work you do not want to be marked.**
- **In all calculations, show clearly how you work out your answer.**

INFORMATION

- **The maximum mark for this paper is 70.**
- **The marks for questions are shown in brackets.**
- **You are expected to use a calculator where appropriate.**
- **You are reminded of the need for good English and clear presentation in your answers.**

DO NOT TURN OVER UNTIL TOLD TO DO SO



0	1
---	---

This question is about acids and bases.

0	1	.	1
---	---	---	---

What is the pH of sulfuric acid? [1 mark]

Tick (✓) ONE box.

1

7

14

0	1	.	2
---	---	---	---

An acid reacts with zinc to produce zinc chloride and hydrogen.

Which acid reacts with zinc to produce zinc chloride? [1 mark]

Tick (✓) ONE box.

Hydrochloric acid

Nitric acid

Sulfuric acid

[Turn over]



01.3

**What type of substance is zinc chloride?
[1 mark]**

Tick (✓) ONE box.

Alkali

Base

Salt

01.4

An alkali is a base in solution.

Which compound is an alkali? [1 mark]

Tick (✓) ONE box.

Sodium hydroxide

Sodium nitrate

Sodium sulfate

[Turn over]



0	1	.	5
---	---	---	---

The formula of the copper ion is Cu^{2+}

The formula of the oxide ion is O^{2-}

What is the formula of copper oxide?
[1 mark]

Tick (✓) ONE box.

Cu_2O_2

CuO_2

CuO

A student reacts an acid with copper oxide.

0 1 . 6

The reaction between the acid and copper oxide is very slow at room temperature.

How could the student speed up the reaction? [1 mark]

[Turn over]

0	1	.	7
---	---	---	---

Complete the sentence to show how the student makes sure that ALL the acid reacts.

Choose the answer from the list. [1 mark]

- **in excess**
- **in solution**
- **molten**
- **soluble**

The student adds copper oxide to the acid until the copper oxide is

_____.

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[Turn over]

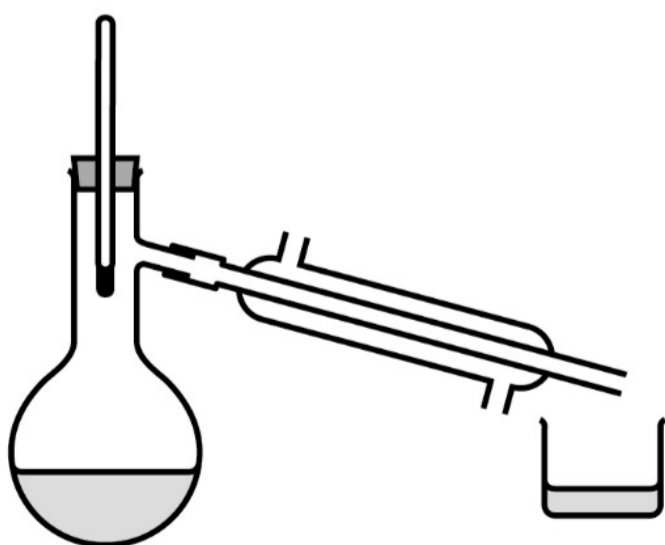
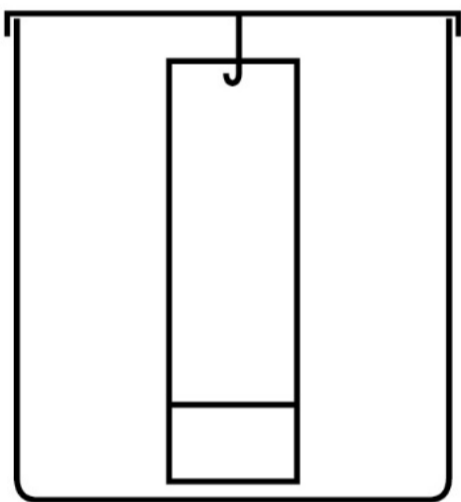


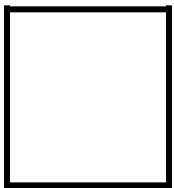
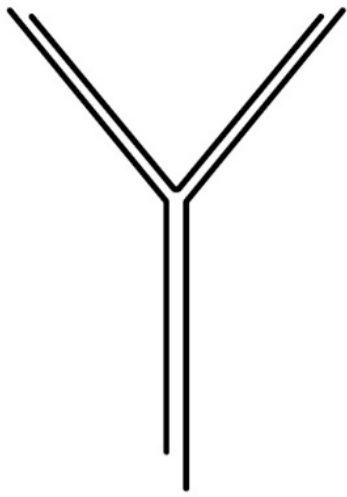
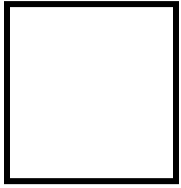
01.8

The student filters the unreacted copper oxide from the solution.

Which apparatus does the student use?
[1 mark]

Tick (✓) ONE box.





0 1 . 9

What process is used to produce crystals of a salt from a salt solution? [1 mark]

[Turn over]

9



0	2
---	---

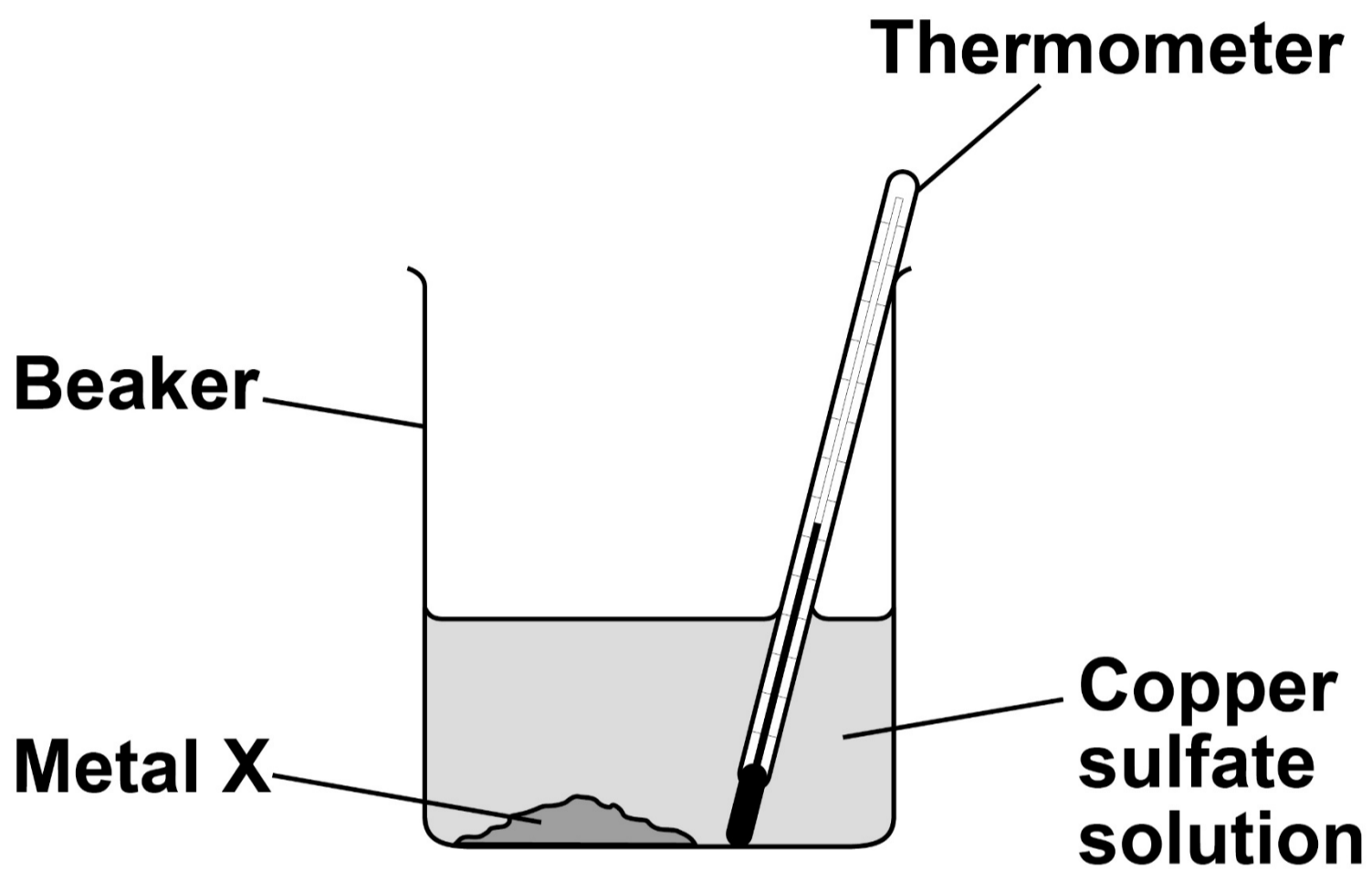
A student investigated the temperature change when metal X was added to copper sulfate solution.

This is the method used.

- 1. Add 25 cm³ of copper sulfate solution to a beaker.**
- 2. Measure the temperature of the copper sulfate solution.**
- 3. Add 1.0 g of metal X and stir.**
- 4. Measure the highest temperature reached when metal X is added to copper sulfate solution.**
- 5. Repeat steps 1 to 4 with different metals.**

FIGURE 1 shows the apparatus used.

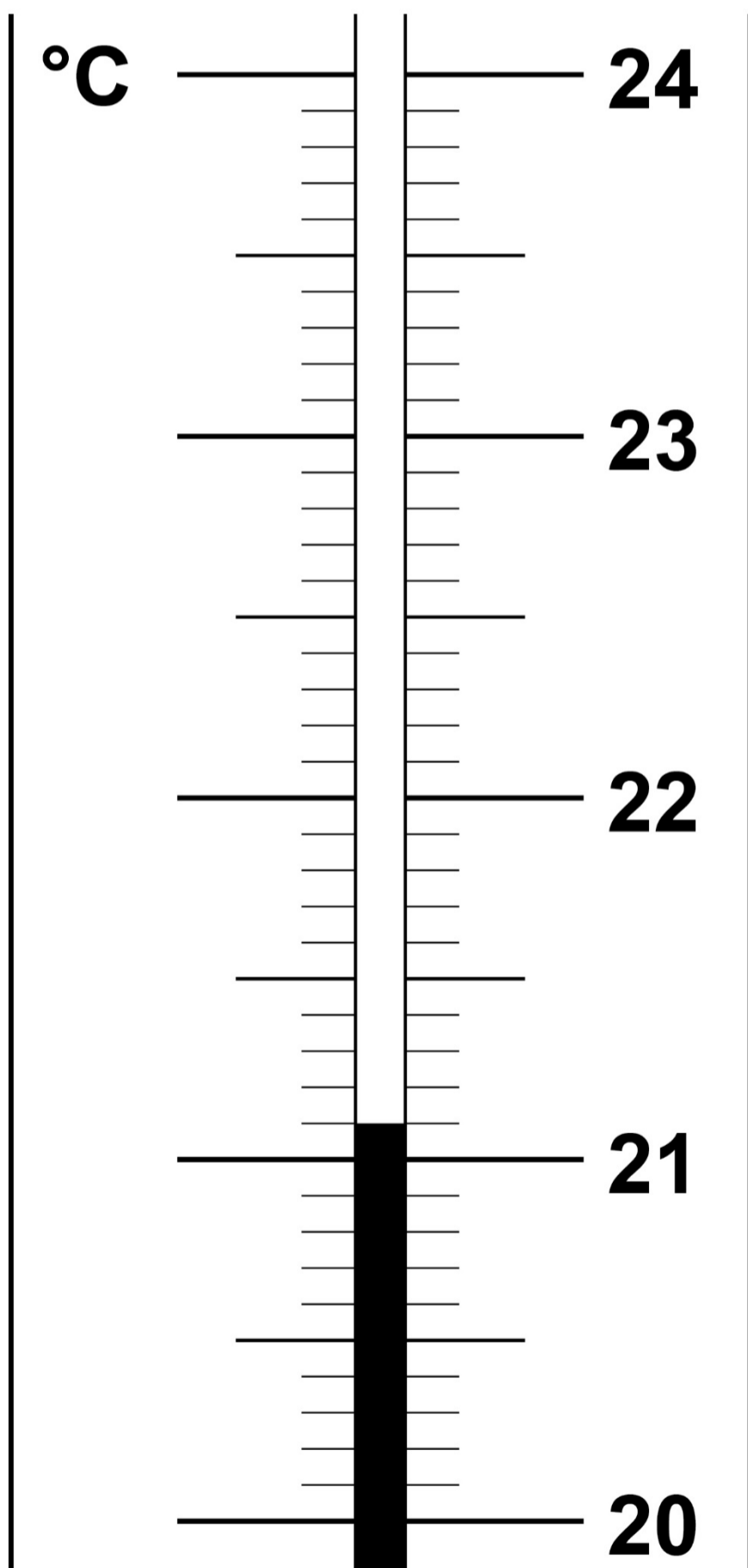
FIGURE 1



[Turn over]

FIGURE 2 shows the thermometer reading of the copper sulfate solution at the start of the investigation.

FIGURE 2



02.1

The highest temperature reached when metal X was added to copper sulfate solution was 35.5 °C

Determine the temperature change when metal X is added to copper sulfate solution.

Use FIGURE 2, on page 16. [2 marks]

Highest temperature = 35.5 °C

Temperature at start = _____ °C

Temperature change = _____ °C

[Turn over]



0 2 . 2

**Give TWO variables the student should keep the same in this investigation.
[2 marks]**

1 _____

2 _____

0 2 . 3

The student repeated the experiment with metal Y.

TABLE 1, on the opposite page, shows four results for metal Y.

TABLE 1

	Test 1	Test 2	Test 3	Test 4
Temperature change in °C	9.2	7.3	9.5	9.2

Calculate the mean temperature change for metal Y.

Do NOT include the anomalous result in your calculation. [2 marks]

Mean temperature change = _____ °C

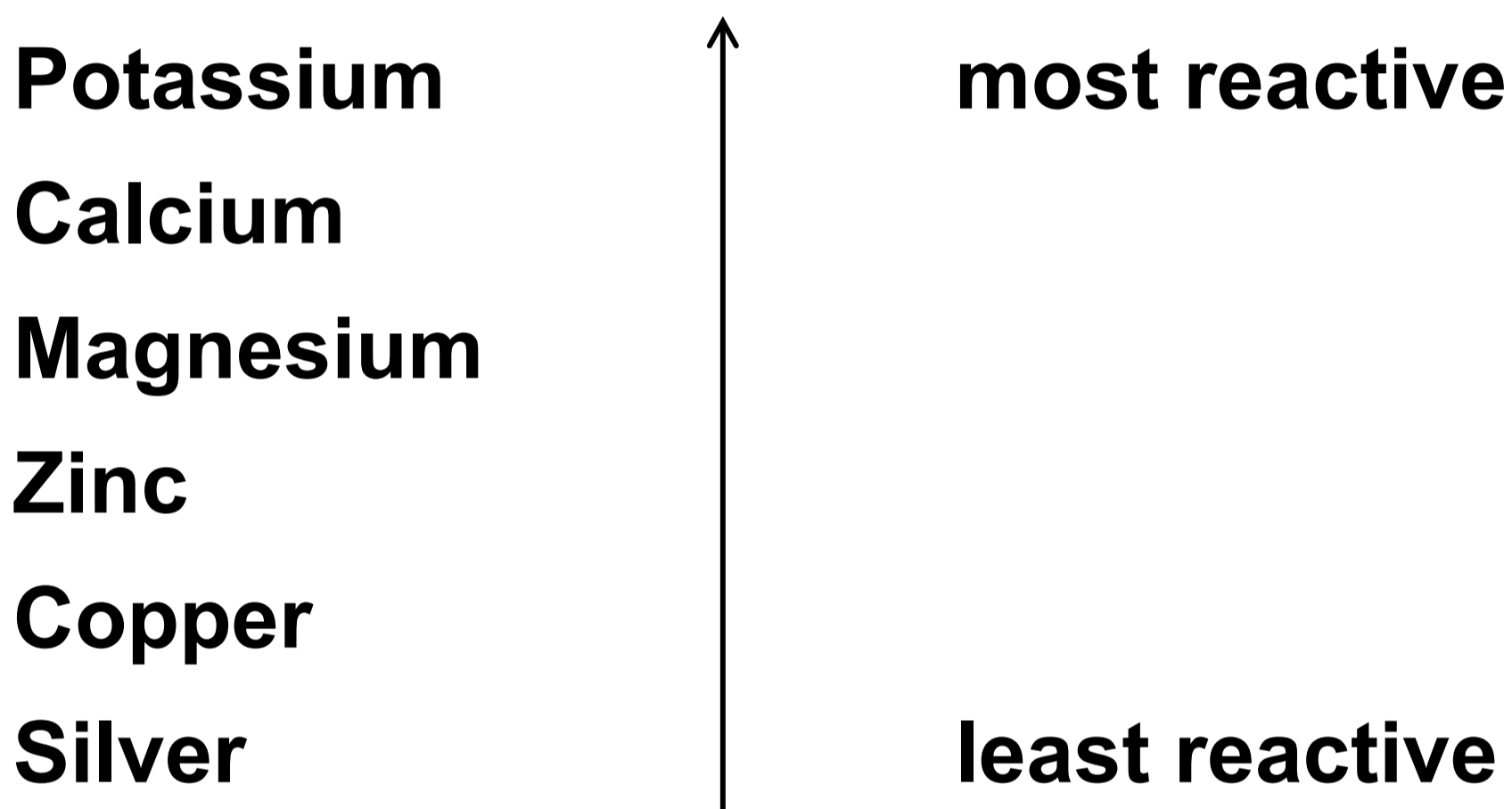
[Turn over]



The more reactive the metal added to copper sulfate solution, the greater the temperature change.

FIGURE 3 shows a reactivity series.

FIGURE 3



02.4

The student repeated the experiment.

The student added:

- magnesium to copper sulfate solution
- an unknown metal A to copper sulfate solution.



TABLE 2 shows the results.

TABLE 2

Metal	Temperature change in °C
Magnesium	12
Metal A	8

The student concludes metal A is zinc.

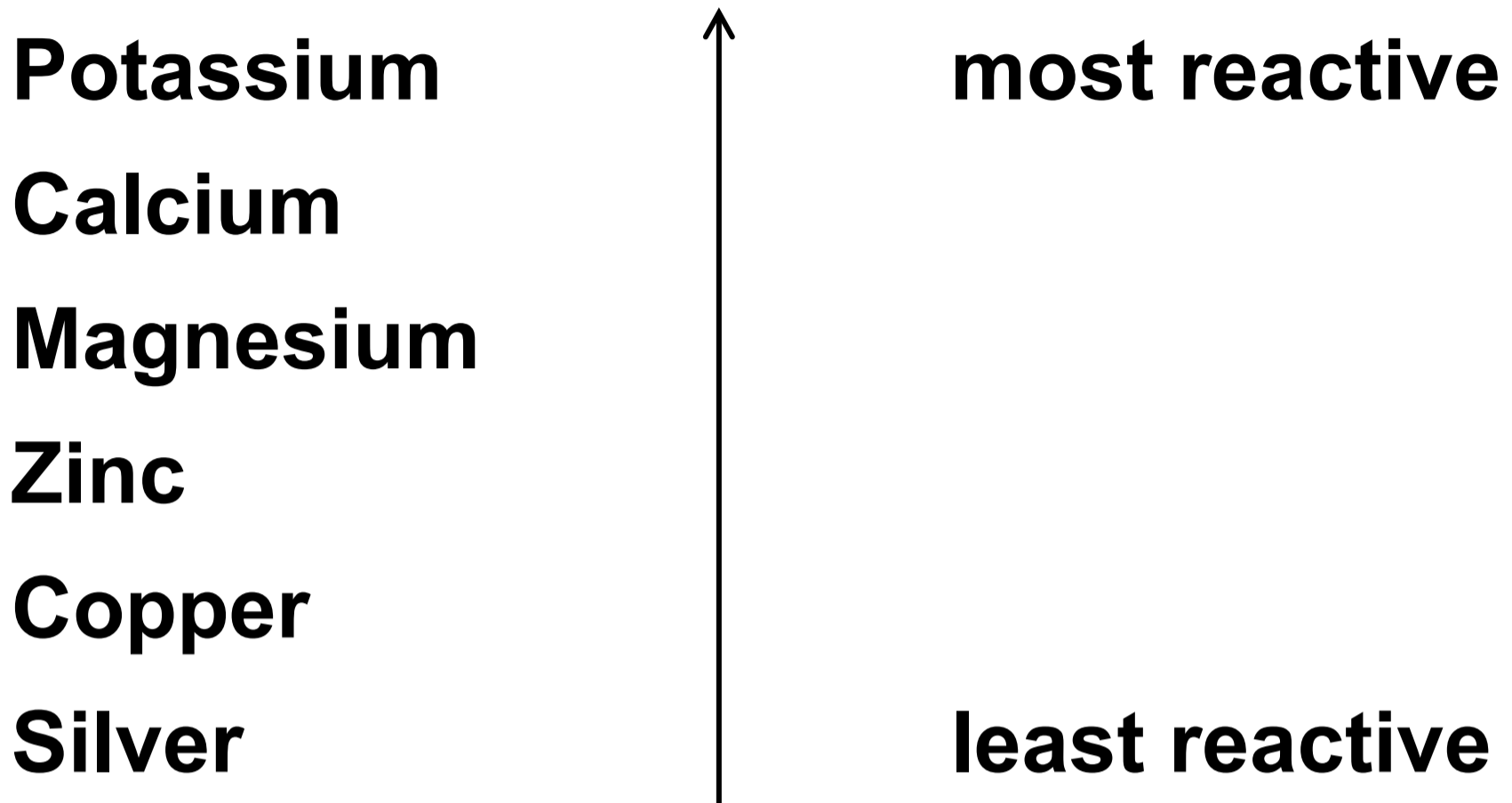
Give ONE reason why the student is correct.

**Use FIGURE 3, on page 20, and TABLE 2.
[1 mark]**

[Turn over]



Repeat of FIGURE 3



0	2	.	5
---	---	---	---

The student did the experiment with silver and copper sulfate solution.

What happens to the temperature of the mixture? [1 mark]

Use FIGURE 3.

Tick (✓) ONE box.

Decreases

Increases

Stays the same

[Turn over]



0	2	.	6
---	---	---	---

Suggest ONE reason why the student should NOT add potassium metal to copper sulfate solution. [1 mark]

0	2	.	7
---	---	---	---

100 cm³ of the copper sulfate solution contains 1.8 g of copper sulfate.

**Calculate the mass of copper sulfate in 25 cm³ of this copper sulfate solution.
[2 marks]**

Mass = _____ g

[Turn over]

11



0	3
---	---

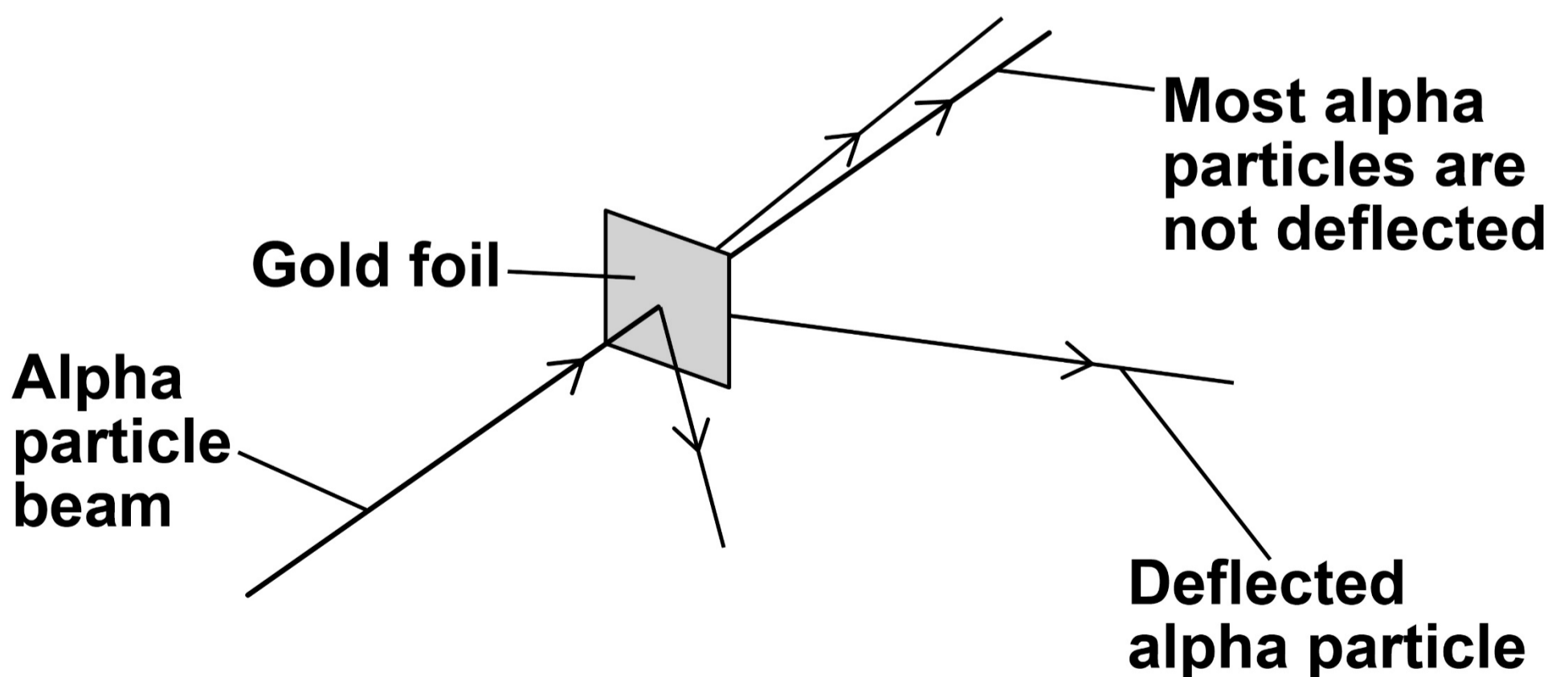
This question is about gold and compounds of gold.

In the alpha particle scattering experiment alpha particles are fired at gold foil.

Alpha particles are positively charged.

FIGURE 4 shows the results.

FIGURE 4



0	3	.	1
---	---	---	---

Some alpha particles are deflected.

Complete the sentence.

Choose the answer from the list. [1 mark]

- **negatively charged**
- **not charged**
- **positively charged**

**Some alpha particles are deflected
because the nucleus of the atom is**

[Turn over]



0	3	.	2
---	---	---	---

Why are most alpha particles NOT deflected? [1 mark]

Tick (✓) ONE box.

The atom is a tiny sphere that cannot be divided.

The atom is mainly empty space.

The electrons orbit the nucleus at specific distances.

0	3	.	3
---	---	---	---

What was ONE conclusion from the alpha particle scattering experiment? [1 mark]

Tick (✓) ONE box.

The mass is concentrated at the centre of the atom.

The mass is concentrated at the edge of the atom.

The mass is spread evenly throughout the atom.

[Turn over]

Gold reacts with the elements in Group 7 of the periodic table.

03.4

**What are Group 7 elements known as?
[1 mark]**

Tick (✓) ONE box.

Alkali metals

Halogens

Noble gases

0	3	.	5
---	---	---	---

Fluorine, chlorine and bromine react with gold.

Which element will be the most reactive with gold? [1 mark]

Tick (✓) ONE box.

Fluorine

Chlorine

Bromine

[Turn over]



0	3	.	6
---	---	---	---

3.94 g of gold reacts with chlorine to produce 6.07 g of gold chloride.

The word equation for the reaction is:

gold + chlorine \longrightarrow gold chloride

Calculate the mass of chlorine that reacts with 3.94 g of gold. [1 mark]

Mass = _____ g

0	3	.	7
---	---	---	---

Calculate the relative formula mass (M_r) of gold chloride (AuCl_3).

Relative atomic masses (A_r):

Cl = 35.5 Au = 197

[2 marks]

Relative formula mass (M_r) = _____

[Turn over]

8



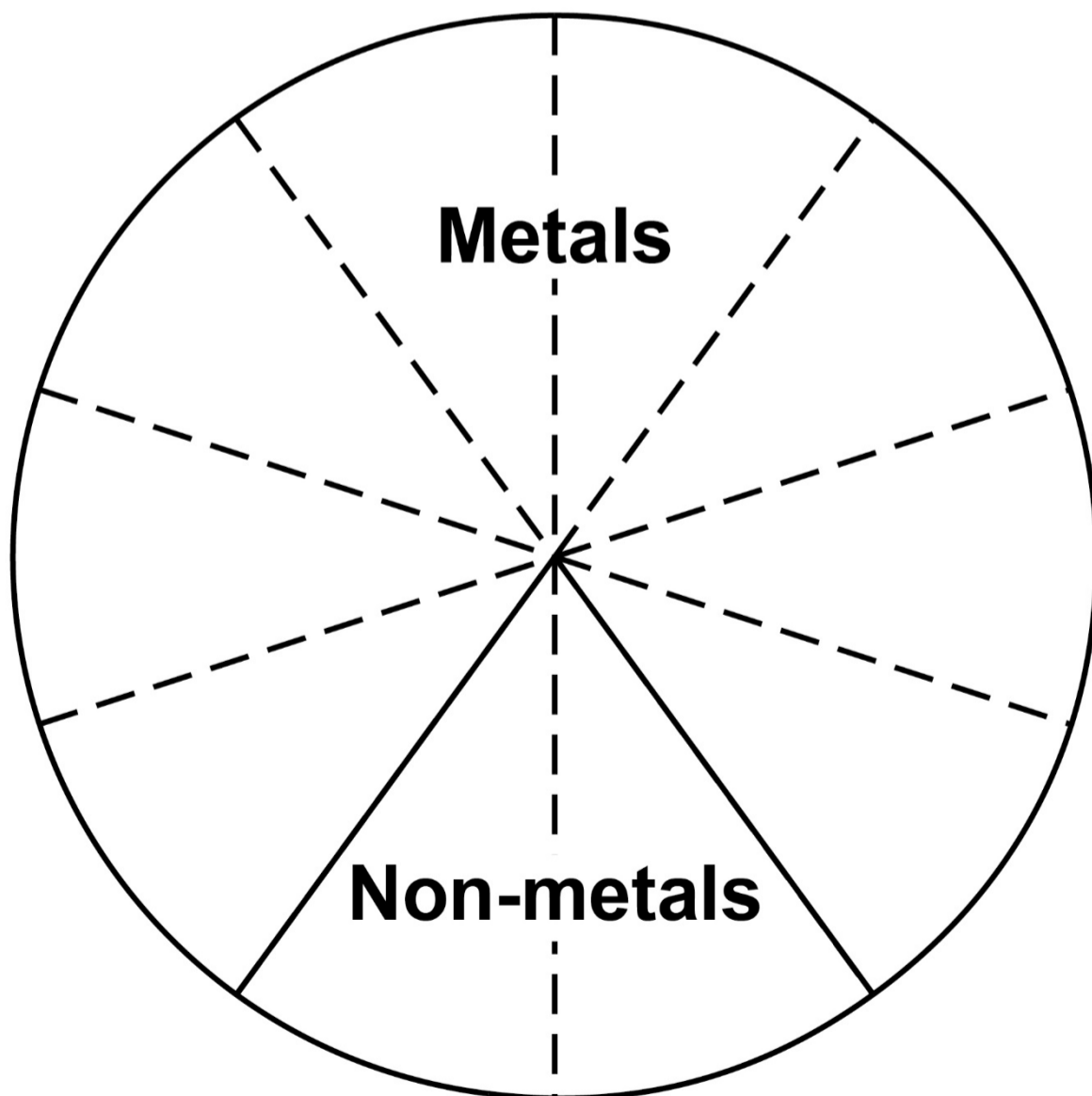
04

This question is about elements and compounds.

04.1

FIGURE 5 shows the proportion of elements in the periodic table that are metals and non-metals.

FIGURE 5



Determine the percentage of the elements in FIGURE 5 that are metals. [2 marks]

Percentage = _____ %

[Turn over]

04.2

**Give TWO physical properties of metals.
[2 marks]**

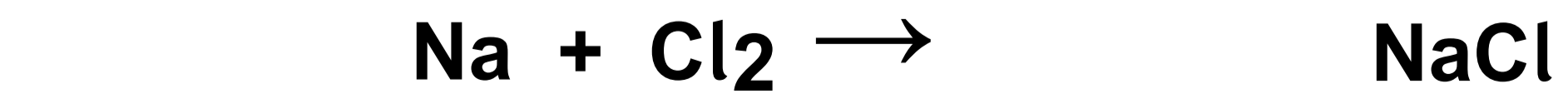
1 _____

2 _____

04.3

Sodium reacts with chlorine to produce sodium chloride.

**Balance the equation for the reaction.
[1 mark]**

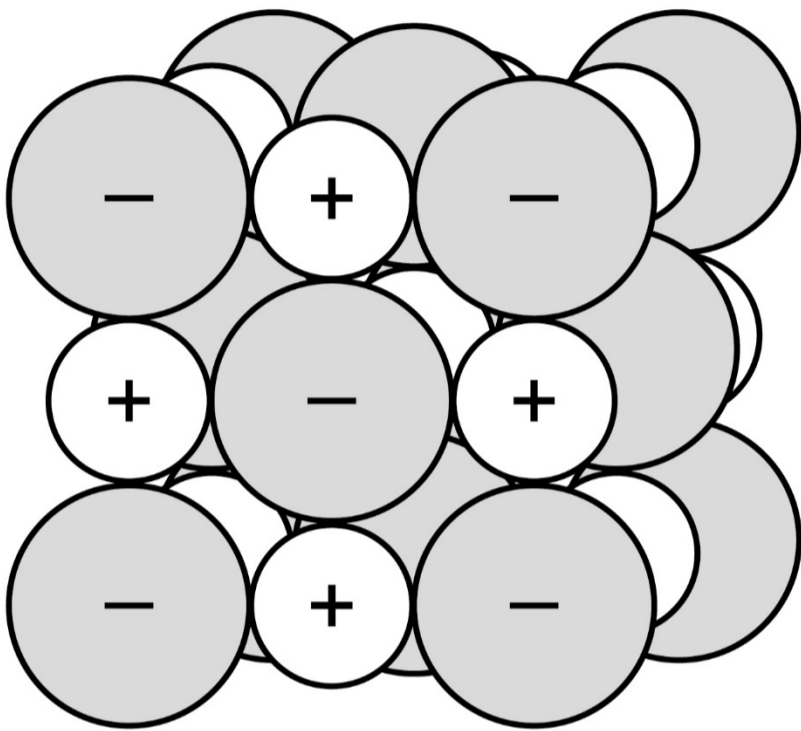


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[Turn over]

FIGURE 6 shows part of the structure of sodium chloride (NaCl).

FIGURE 6



Sodium chloride

0	4	.	4
---	---	---	---

What holds the particles together in sodium chloride?

Use FIGURE 6. [1 mark]

Tick (✓) ONE box.

Electrostatic attractions

Intermolecular forces

Metallic bonds

[Turn over]

0	4	.	5
---	---	---	---

Solid sodium chloride does not conduct electricity.

Give TWO ways in which sodium chloride can be made to conduct electricity.

[2 marks]

1 _____

2 _____

8

0 5

This question is about elements in the periodic table.

0 5 . 1

**What property was used to arrange elements in early periodic tables?
[1 mark]**

Tick (✓) ONE box.

Atomic number

Atomic weight

Mass number

[Turn over]



05.2

In early periodic tables, iodine (I) was placed before tellurium (Te).

Mendeleev placed iodine after tellurium.

FIGURE 7 shows part of Mendeleev's periodic table.

FIGURE 7

16 O	19 F
32 S	35.5 Cl
79 Se	80 Br
128 Te	127 I

Suggest ONE reason why Mendeleev placed iodine in the column shown in FIGURE 7. [1 mark]

[Turn over]

TABLE 3 shows the melting points of three Group 1 metals.

TABLE 3

Metal	Melting point in °C
Lithium	180
Sodium	98
Potassium	63



05.3**What state is lithium at 100 °C?****Use TABLE 3. [1 mark]****Tick (✓) ONE box.****Gas****Liquid****Solid****[Turn over]**

Repeat of TABLE 3

Metal	Melting point in °C
Lithium	180
Sodium	98
Potassium	63

0 5 . 4

Complete the graph in FIGURE 8 on the opposite page.

Use TABLE 3.

You should:

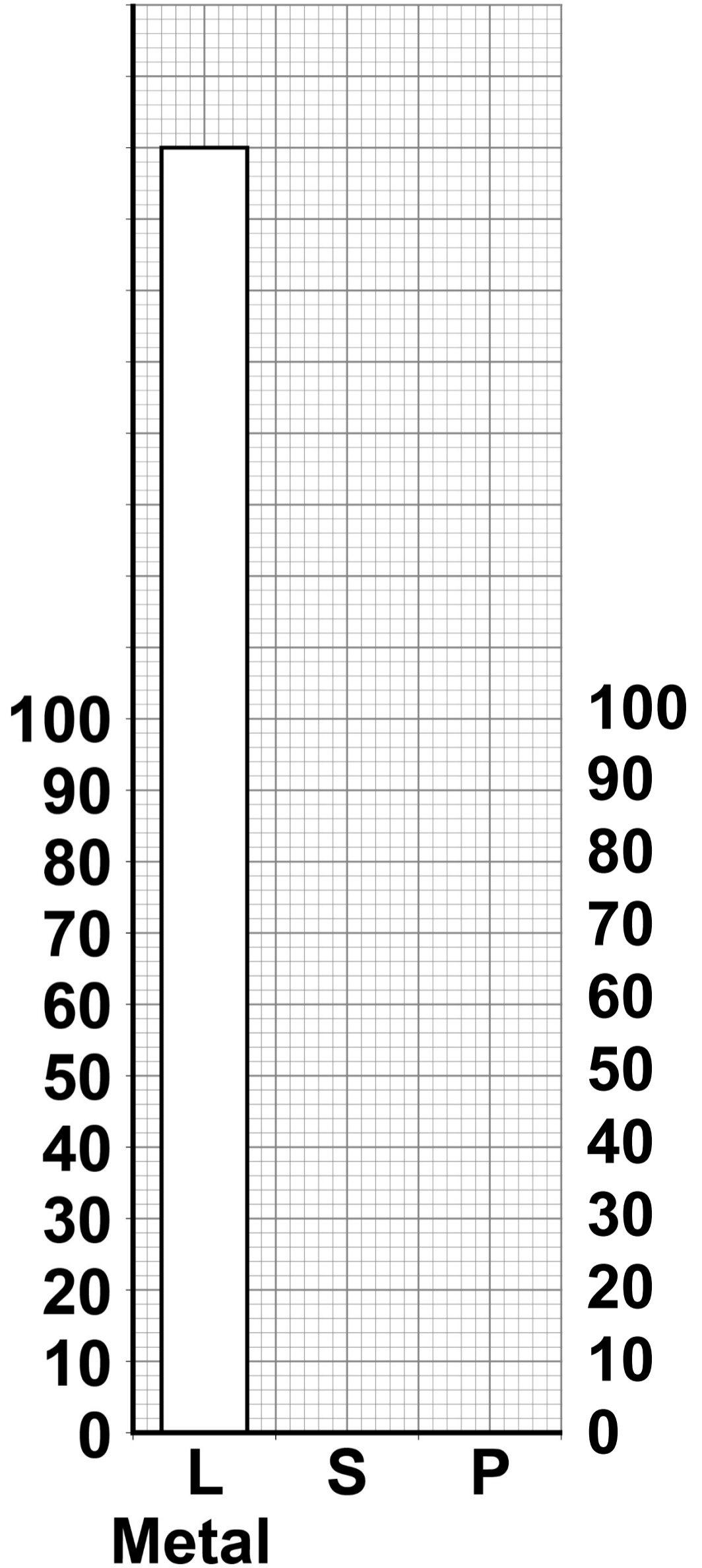
- complete the scale on the y-axis**
- draw bars to show the melting points of sodium and potassium.**

[3 marks]



FIGURE 8

**Melting point
in °C**



KEY

L = Lithium

S = Sodium

P = Potassium



[Turn over]

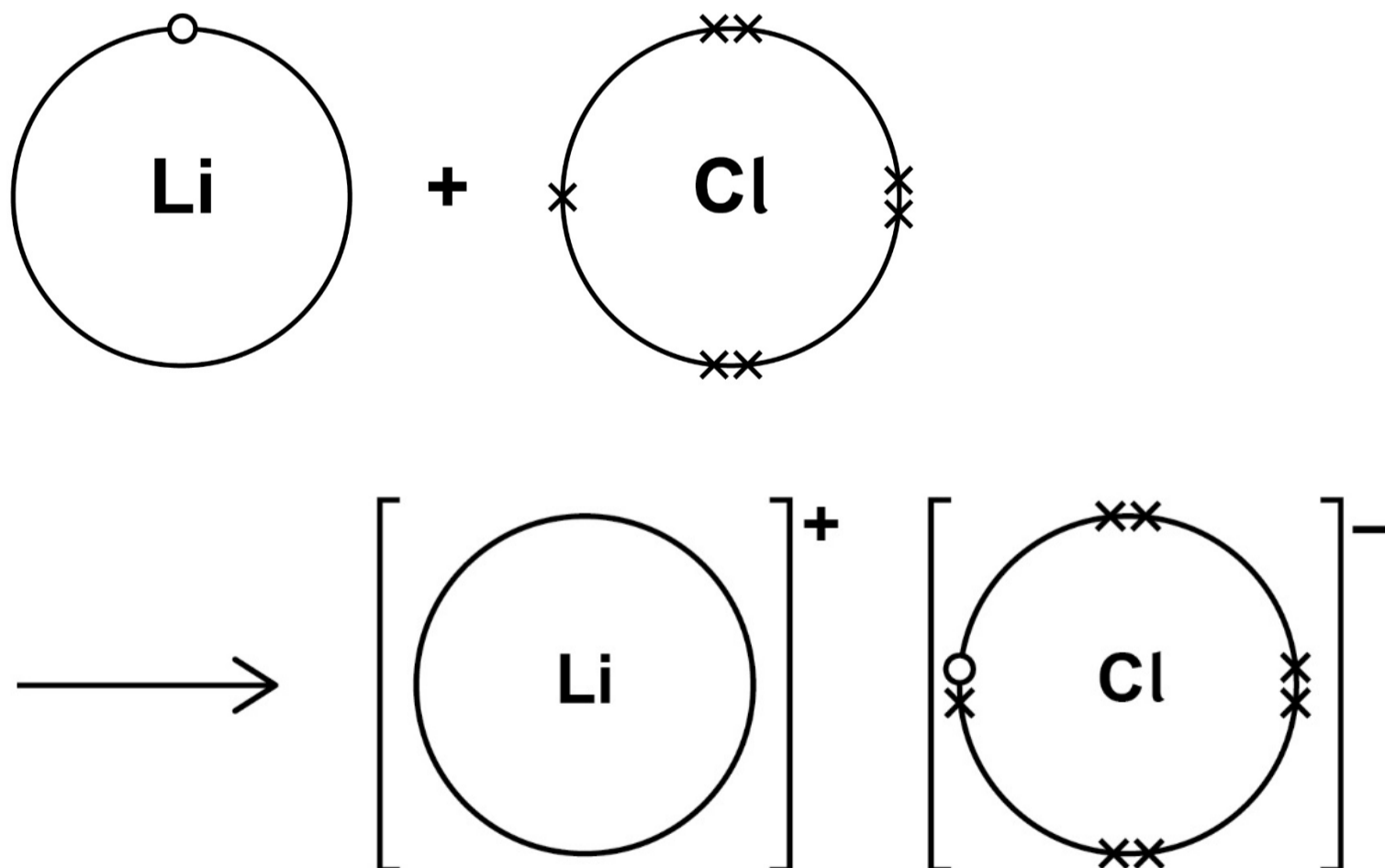
05.5

Lithium reacts with chlorine to produce lithium chloride.

FIGURE 9 shows what happens to the electrons in the outer shells when a lithium atom reacts with a chlorine atom.

The dots (o) and crosses (x) represent electrons.

FIGURE 9



Describe what happens to a lithium atom and to a chlorine atom when they react.

Use FIGURE 9 to answer in terms of electrons. [3 marks]

[Turn over]

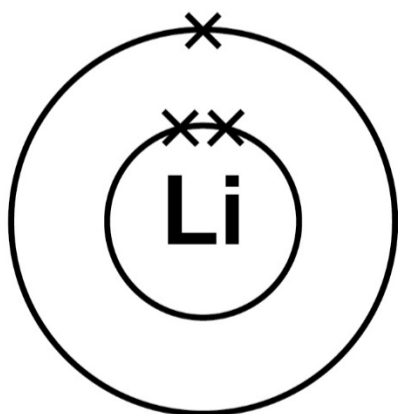
05.6

Lithium and potassium are in the same group of the periodic table.

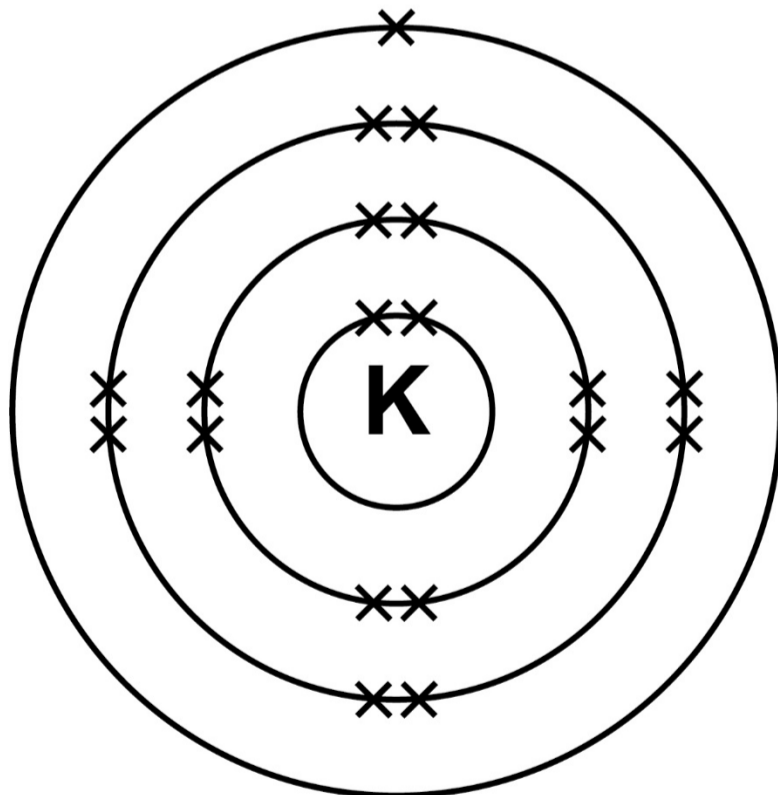
FIGURE 10 represents the electronic structures of a lithium atom and of a potassium atom.

FIGURE 10

Lithium atom



Potassium atom



Give TWO reasons why potassium is more reactive than lithium. [2 marks]

1 _____

2 _____

[Turn over]

11

06

This question is about the extraction of aluminium.

06.1

An aluminium atom is represented as:



Give the number of electrons and neutrons in the aluminium atom.

[2 marks]

Number of electrons _____

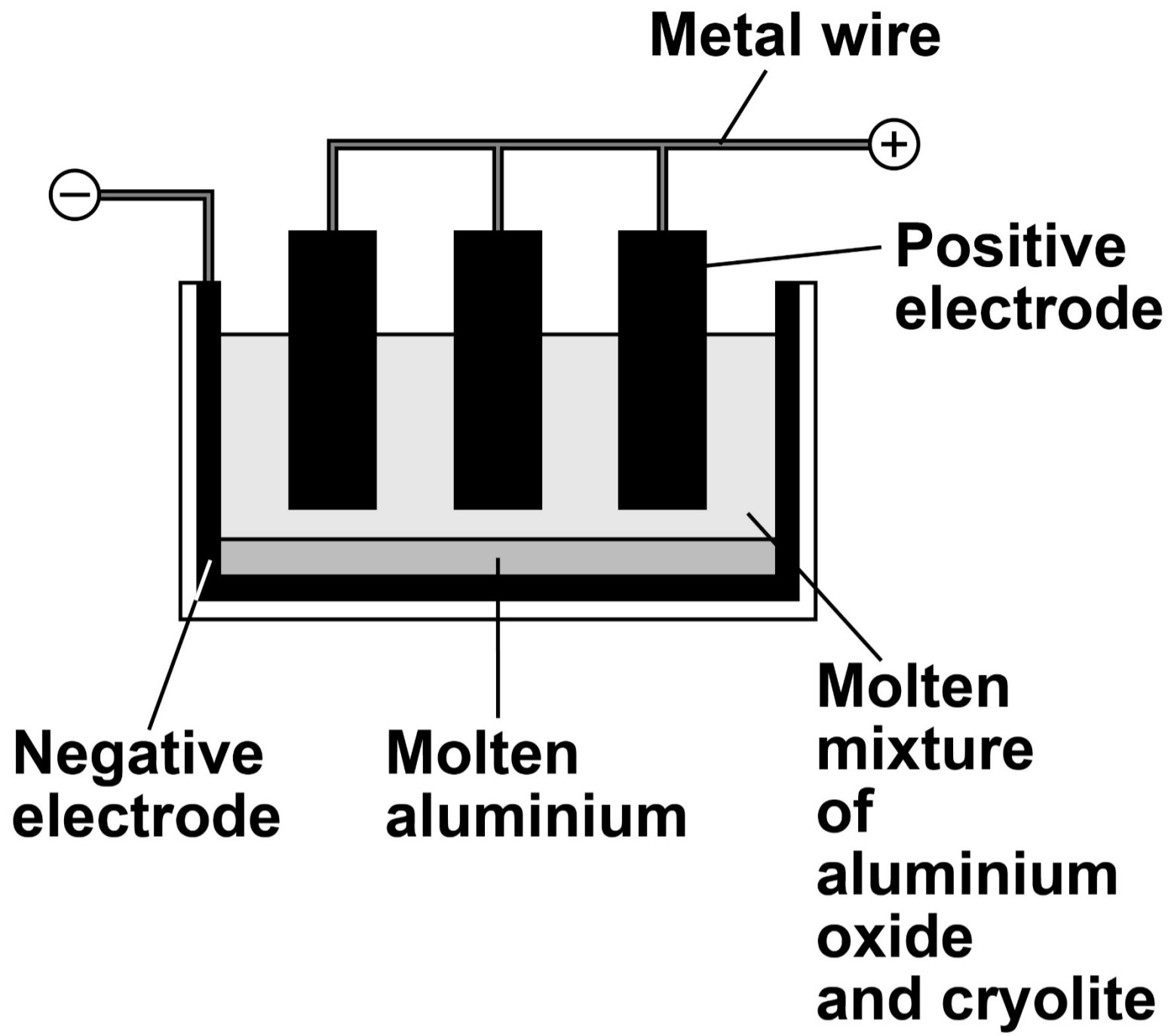
Number of neutrons _____

Aluminium is extracted by the electrolysis of a molten mixture of aluminium oxide and cryolite.

FIGURE 11, on the opposite page, shows the cell used for the electrolysis.



FIGURE 11



[Turn over]



06.2

Aluminium is produced by the reduction of aluminium oxide (Al_2O_3).

**What is meant by the term reduction?
[1 mark]**

0 | 6 | . | 3

Oxygen is formed at the positive carbon electrodes.

Explain why the positive carbon electrodes must be continually replaced.
[3 marks]

[Turn over]



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06.4

A substance conducts electricity because of free moving, charged particles.

What are the free moving, charged particles in a:

- **carbon electrode (made from graphite)**
- **molten mixture of aluminium oxide and cryolite**
- **metal wire?**

[3 marks]

Carbon electrode (made from graphite)

Molten mixture of aluminium oxide and cryolite _____

Metal wire _____

[Turn over]



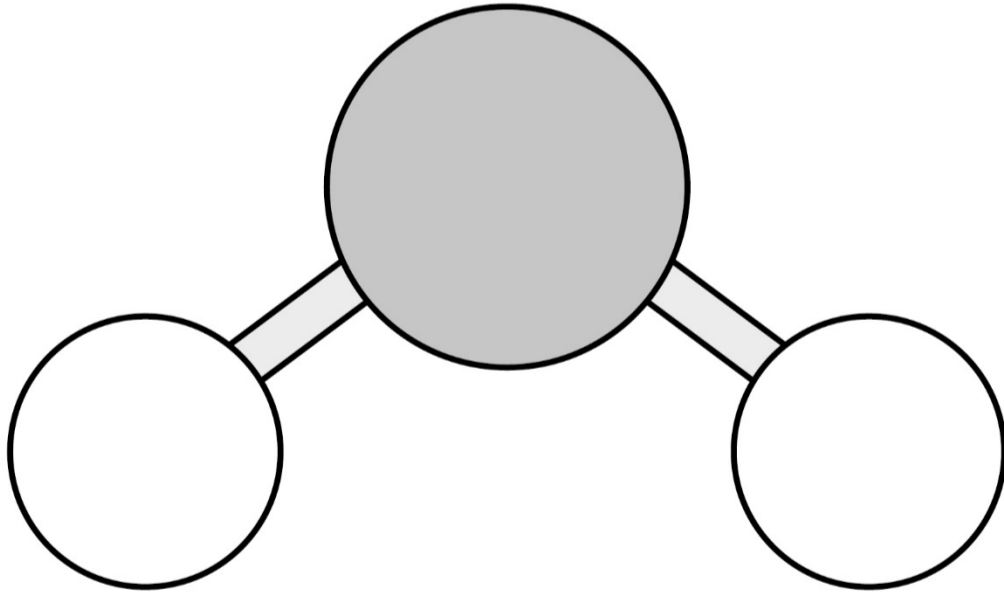
07

This question is about substances with covalent bonding.

07.1

FIGURE 12 shows a ball and stick model of a water molecule (H_2O).

FIGURE 12



Suggest ONE limitation of using a ball and stick model for a water molecule. [1 mark]

07.2

Ice has a low melting point.

Water molecules in ice are held together by intermolecular forces.

Complete the sentence. [1 mark]

Ice has a low melting point because the intermolecular forces are

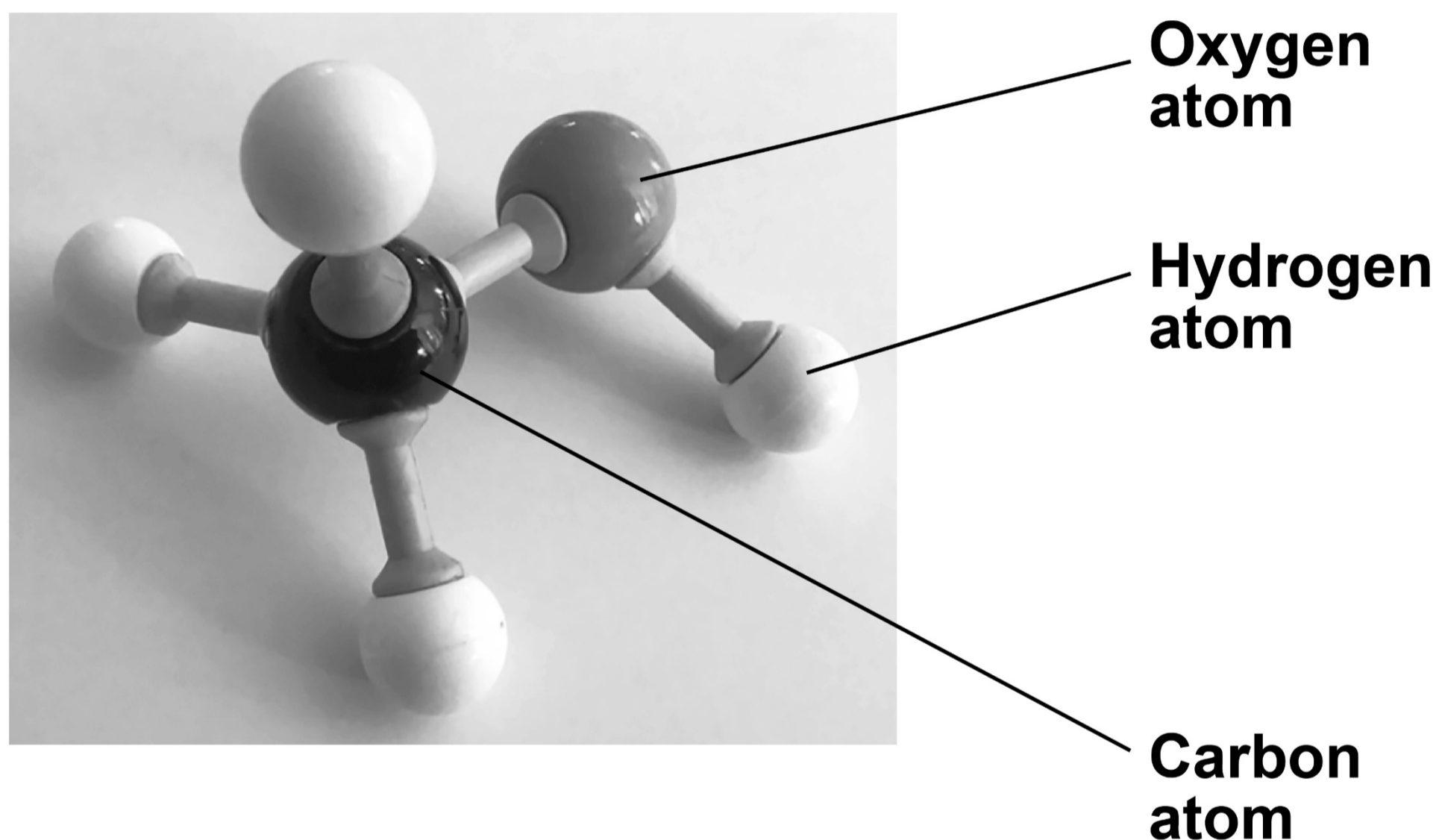
[Turn over]



07.3

FIGURE 13 shows the structure of a molecule.

FIGURE 13



What is the molecular formula of the molecule in FIGURE 13? [1 mark]

Diamond has a giant covalent structure.

07.4

What is the number of bonds formed by each carbon atom in diamond? [1 mark]

Tick (✓) ONE box.

2

3

4

8

[Turn over]



0	7	.	5
---	---	---	---

Give TWO physical properties of diamond. [2 marks]

1 _____

2 _____

0	7	.	6
---	---	---	---

Name two other substances with giant covalent structures. [2 marks]

1 _____

2 _____

8



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[Turn over]

0	8
---	---

Some students investigated the thermal decomposition of metal carbonates.

The word equation for the reaction is:

metal carbonate

→ metal oxide + carbon dioxide

The students made the following hypothesis:

‘When heated the same mass of any metal carbonate produces the same mass of carbon dioxide.’

The students heated a test tube containing copper carbonate.

TABLE 4, on the opposite page, shows their results.



TABLE 4

Time the test tube containing copper carbonate was heated in mins	0	2	4	6
Mass of test tube and contents in g	17.7	17.1	17.0	17.0

Plan a method the students could use to test their hypothesis.

You should show how the students use their results to test the hypothesis.

You do NOT need to write about safety precautions. [6 marks]

[Turn over]



END OF QUESTIONS



<hr/> 6

Additional page, if required.

Write the question numbers in the left-hand margin.

Additional page, if required.

Write the question numbers in the left-hand margin.

Lined writing area with 16 horizontal lines and a vertical margin line on the left.

Additional page, if required.

Write the question numbers in the left-hand margin.



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For Examiner's Use	
Question	Mark
1	
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8	
TOTAL	

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