

**MARK SCHEME for the October/November 2011 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/35

Paper 3 (Advanced Practical Skills 1),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

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Question	Sections	Indicative material	Mark	
1 (a)	PDO Layout Recording	I Pairs of thermometer readings and time unambiguously recorded. <i>Minimum of three pairs.</i>	1	[5]
		II Correct headings and units. <i>Units must have solidus: /s; brackets: (s); or describe in words: time in seconds or time in s, solidus/°C; or brackets (°C); or describe in words: temperature in °C.</i> <i>No repeats of unit in table to individual readings.</i>	1	
		III Time recorded to 1 second and temperature to 0.5°C. (Must have at least one at 0.5°C.)	1	
	MMO Decisions	IV Five (minimum) different experiments carried out.	1	
		V Initial temperatures span the range specified in the question. At least 1 at or below 40°C, at least 1 above 50°C and no two within 3°C (minimum 3 readings). If more than 5 readings can be within 3°C.	1	
(b)	ACE Interpretation	I Correct means and rates for highest 2 temperatures and lowest 2 temperatures. <i>Use candidate's times (not corrected).</i>	1	[6]
	PDO Display MMO Quality	II 1000/time recorded 3–4sf.	1	
Calculate log rate for chosen temperatures. Calculate $\frac{\log \text{rate } 2 - \log \text{rate } 3}{T_2 - T_3}$ to 3 significant figures (factor A). Calculate $(T_1 - T_2) \times \text{factor A}$ and add this to log rate 2. Compare this value with candidate's log rate for T1 and calculate δ . If $\delta > 0.05$ but ≤ 0.10 award III , If $\delta \leq 0.05$ award III and IV . Repeat for $(T_3 - T_4) \times \text{factor A}$ and subtract from log rate T3. If $\delta > 0.05$ but ≤ 0.10 award V , If $\delta \leq 0.05$ award V and VI . If 3 experiments, use slowest 2 for 'standard' and award Q marks for fastest (maximum 2)			1 1 1 1	

(c)	PDO Layout	(i) I All points plotted to use at least 5 large squares on vertical axis and 4 on horizontal axis including position of 20°C.	1	[4]
	ACE Interpretation	II x-axis to allow extrapolation to 20°C. III An appropriate line of best fit is drawn. (ii) IV Correct value to 0.5°C or 1 dp of 1000/t from graph (ignore units).	1 1 1	
(d)	PDO Display	Uses temperature values 10 apart from graph and quotes rates/chooses rates that are doubled and quotes temperatures.	1	[2]
	ACE Conclusions	Relevant comment on data made. This can come from experimental results.	1	
(e)	ACE Interpretation	(i) Fastest reaction/first reaction.	1	[2]
		(ii) Expression for % error ecf from (i).	1	
(f)	ACE Interpretation	Temperature change is not the same for each run of the mixture/ FA 2 not at the same temperature as FA 1 before mixing/difficulty of gauging same level of colour/cannot start clock and pour solutions at same time/reusing boiling tubes could affect concentration. <i>Not: human error/heat loss or gain/human reaction time.</i>	1	[1]
(g)	ACE Improvements	Use of thermostatically controlled water bath/data logger for colour intensity/colorimeter/get help from another student. <i>Improvement must correspond to error specified.</i> <i>Not: automatic timer.</i>	1	[1]
(h)	ACE Improvements	Same volume FA 2 .	1	[3]
		Change volume of FA 1 and keep total volume constant by adding water for several volumes of FA 1 .	1	
		All experiments carried out at the same temperature.	1	
				[Total: 24]

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FA 3 is BaCl ₂ or Ba(NO ₃) ₂ ; FA 4 is H ₂ SO ₄ ; FA 5 is NH ₄ Cl + Na ₂ SO ₃					
2 (a)	MMO Decisions	(i) I	Selects named reagent involving CrO ₄ ²⁻ or CO ₃ ²⁻ (solution) or magnesium.	1	[5]
	PDO Layout	(ii) II	Tabulates evidence of three tests carried out with no repeat headings (irrespective of reagents).	1	
	MMO Collection	III	FA 3 yellow ppt or white ppt or no change.	1	
		IV	FA 4 (yellow solution turns) orange or effervescence or effervescence.	1	
		V	FA 5 yellow solution/no reaction/no reaction.	1	
<p><i>Do not allow NaOH for I but allow observations to include T rise for FA 4. If acid as reagent can score only II. Acidified potassium dichromate is 1 reagent. Do not credit as reagent but credit all observations FA 3 yellow ppt, FA 4 no change, FA 5 green.</i></p>					
(b)	MMO Collection	I	FA 3 + FA 4 white ppt.	1	[4]
		II	FA 4 + FA 5 no reaction or slow effervescence.	1	
		III	FA 5 + FA 3 white ppt.	1	
		IV	ppt insoluble in HCl in (i) and soluble in HCl in (iii).	1	
(c)	MMO Collection		<p>Mark vertical columns</p> <p>FA 3 no ppt/ignore faint/slight white ppt.</p> <p>FA 5 no ppt and gas/ammonia turning red litmus blue on warming.</p> <p>FA 4 no visible reaction.</p> <p>Cross out any identified as acid in (a) (iii).</p>	2	[2]

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(d)	ACE Conclusions	<p>All conclusions must follow observations. For each unknown. One mark for ion and one mark for satisfactory evidence.</p> <p>FA 3 must be Ba^{2+} or Ca^{2+} to gain credit.</p> <p>FA 3 if CrO_4^{2-} in (a) (i), Ba^{2+} (1).</p> <p>Evidence: yellow ppt or white ppt with FA 4/H_2SO_4 and no ppt/(faint) white ppt with NaOH (1) (Must have 2 pieces of evidence.)</p> <p>If CrO_4^{2-} not used in (a) (i) Ba^{2+} and/or Ca^{2+} (1).</p> <p>Evidence: faint white/no ppt NaOH and white ppt with FA 4/sulfuric acid/no NH_3 when heated with NaOH (1). (Must have 2 pieces of evidence.)</p> <p>FA 5 NH_4^+ (1).</p> <p>Evidence: formation NH_3 in (c) (1).</p> <p>SO_3^{2-} (1).</p> <p>Evidence: formation SO_2 in (a) or (b) (1).</p>	1 1 1 1	[4]
(e)	MMO Decisions	Cream ppt (not off white) with $AgNO_3$ (partially soluble/insoluble in aq NH_3)	1	[1]
[Total: 16]				