CAMBRIDGE INTERNATIONAL EXAMINATIONS

Cambridge International General Certificate of Secondary Education

MARK SCHEME for the October/November 2015 series

0652 PHYSICAL SCIENCE

0652/21

Paper 2 (Core Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2015 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.



	Cambridge 10001 Cottober 110101	0002 2.
1	(a) atom of an element with same number of proton/proton number; different number of neutrons/nucleon or neutron number;	[1] [1]
	(if no reference to 'number' 1 max, for both proton and nucleon)	
	(b) atomic/proton number; mass/nucleon number;	[1] [1]
	(c) 2 electrons in first level and 4 in outer level;	[1]
		[Total 5]
2	(a) Arrow from C; vertically downwards;	[1] [1]
	(b) (i) <u>Use of mass \times g (= 80 \times 10); 800 N (accept use of (9.8 or 9.81 N kg⁻¹);</u>	[1] [1]
	(ii) Use of weight \times distance (= 800 \times 6.0) (ecf); = 4800 (N m);	[1] [1]
	<pre>(iii) decreases; (moment = force × distance from X) distance decreases owtte;</pre>	[1] [1]
		[Total 8]
3	brass; graphite/sulfur; air; graphite; chlorine;	[1] [1] [1] [1]
		[Total 5]
4	(a) conduction;	[1]
	(b) copper fastest, iron slowest; brass quicker conductor than aluminium;	[1] [1]
		[Total 3]

Mark Scheme
Cambridge IGCSE – October/November 2015

Syllabus 0652 Paper 21

			Cambridge IGCSE – October/November 2015	0652 21	
5	(a)	(i)	loses one/an electron;	[1]	
		(ii)	C <i>l</i> ⁻ ; 18 ;	[1] [1]	
	(b)		lectrons round chlorine ; ared pair between hydrogen and chlorine ;	[1] [1]	
	(c)		lium hydroxide/sodium oxide ; ter ;	[1] [1]	
		OR			
			lium carbonate/hydrogencarbonate ; er AND carbon dioxide ;	[1] [1]	
		(ac	cept correct formulae)		
				[Total 7]	
6	(a)	(i)	I clearly marked equal distance behind the mirror as object is in from with the object;	nt and in line [1]	
			(accept very small angle between incident and reflected ray, < 5°)		
		(ii)	Ray 1 correctly reflected back along its own path;	[1]	
		(iii)	Ray 2 correctly reflected ;	[1]	
		(iv)	normal drawn and 'r' correctly identified;	[1]	
		(v)	Ray 1 correctly continued along its own path; Ray 2 correctly continued along its own path;	[1] [1]	
		(vi)	E at a suitable point with between the rays	[1]	
	(b)	angle of reflection = angle of incidence ;			
	(c)	virt	ual (accept cannot be projected onto a screen);	[1]	
				[Total 9]	
7	(a)	hyc	ogen – 2 ; Irogen – 8 ; Igen – 4 ;	[1] [1] [1]	
			n be listed in any order with the correct number) vard one mark if all three names correct and no other marks gained)		

Mark Scheme

Syllabus

		Cambridge IGCSE – October/November 2015 0652	21			
	(b)	0 + 12 + 48 ; 00 ;	[1] [1]			
		(correct final answer with no working scores 2 marks, ignore any unit given)				
	(c)) any number between 4 and 6.9;	[1]			
		(4 is acceptable but 7 is not)				
) 7;	[1]			
			[Total 7]			
8	(a)	alloons are charged (by rubbing)	[1]			
		(accept charge transferred from jumper to balloon or vice versa)				
		oth have same charge (accept both positive/negative); ke charges repel;	[1] [1]			
	(b)	rater conducts (charge) ; rater removes charge/balloons discharged ;	[1] [1]			
			[Total 5]			
9	(a)	ny one from: cloured ions/compounds/; nore than one ion formed/different oxidation states/variable valencies; seful catalysts/form complexes; igh densities/melting points;	[1]			
		(accept conducts electricity or energy)				
	(b)	rsenic/selenium/bromine/krypton;	[1]			
	(c)) malachite/copper pyrites ;	[1]			
) gold/silver/mercury/platinum;	[1]			
) unreactive ;	[1]			

Mark Scheme

Syllabus

Paper

Pá	age :	5	Mark Scheme	Syllabus	Paper
			Cambridge IGCSE – October/November 2015	0652	21
	(d)	(i)	no reaction/no change/OWTTE; copper formed/iron dissolves/solution turns colourless;		[1] [1]
		(ii)	iron is more reactive than copper;		[1]
					[Total 8]
10	(a)	(i)	heat/energy given out;		[1]
		(ii)	$2C_2H_2 + 5O_2 \rightarrow 4CO_2 + 2H_2O$; (all formulae correct – 1 mark; correct balancing – 1 mark)		[2]
	(b)	(i)	carbon monoxide ;		[1]
		(ii)	poisonous/toxic/prevents transport of oxygen in blood/bonds with haemoglobin;		[1]
	(c)	(i)	members differ by CH ₂ /same general formula/functional group;		[1]
			(accept similar chemical properties/physical properties increase do	own series)	
		(ii)	ethane has carbon-carbon single bond; ethaene has carbon-carbon double bond;		[1] [1]
					[Total 8]
11	(a)		rect circuit diagram for fuse — ; cept		[1]
	(b)	= 3	$\frac{\text{of }}{75}R = V/I(=12/3.2);$ n or Ω ;		[1] [1] [1]
	(c)		; st be greater than 3.2A (accept for 13A fuse); smallest above 3.2A/relevant comment re 13A fuse;		[1] [1] [1]
		(If 3	A is chosen and reason given is that it is the nearest to current allo	w 1c).	
	(d)	(i)	lamp correctly drawn in parallel with the original lamp;		[1]
		(ii)	circuit current/current through fuse now larger; greater than 5 A/= 6.4 A;		[1] [1]

[Total 10]

Page 6		3	Mark Scheme	Syllabus	Paper
			Cambridge IGCSE – October/November 2015	0652	21
12	(a)	ran	domness of decay ;		[1]
	(b)	(i)	$4600 \mathrm{s}^{-1}(\mathrm{Bq})$;		[1]
		(ii)	Indication on graph of finding time at which count rate halves ; 25 $\pm2\text{s}$;		[1] [1]
	(c)	Pro	tective clothing/use tongs/short exposure time/shielding etc.;		[1]

[Total 5]