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International General Certificate of Secondary Education

MARK SCHEME for the May/June 2006 question paper

0653 COMBINED SCIENCE

0653/03

Paper 3, maximum raw mark 80

These mark schemes are published as an aid to teachers and students, to indicate the requirements of the examination. They show the basis on which Examiners were initially instructed to award marks. They do not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published Report on the Examination.

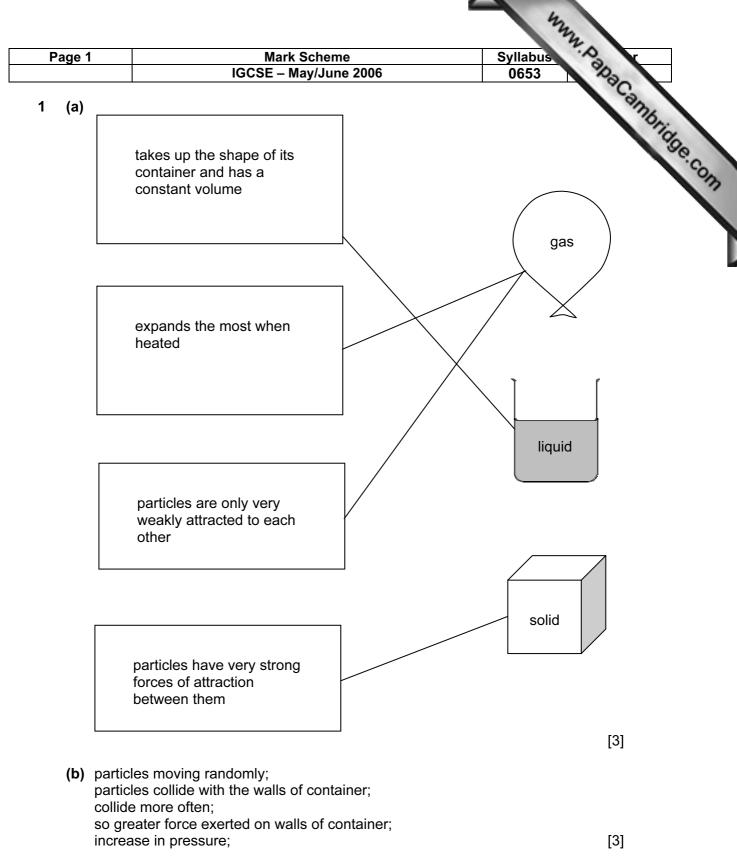
All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the Report on the Examination.

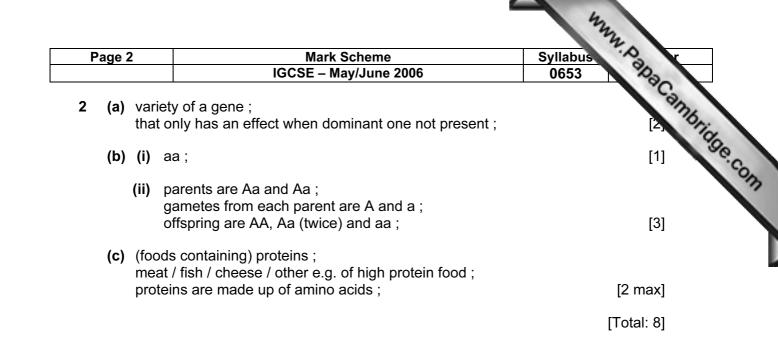
The minimum marks in these components needed for various grades were previously published with these mark schemes, but are now instead included in the Report on the Examination for this session.

CIE will not enter into discussion or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the May/June 2006 guestion papers for most IGCSE and GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



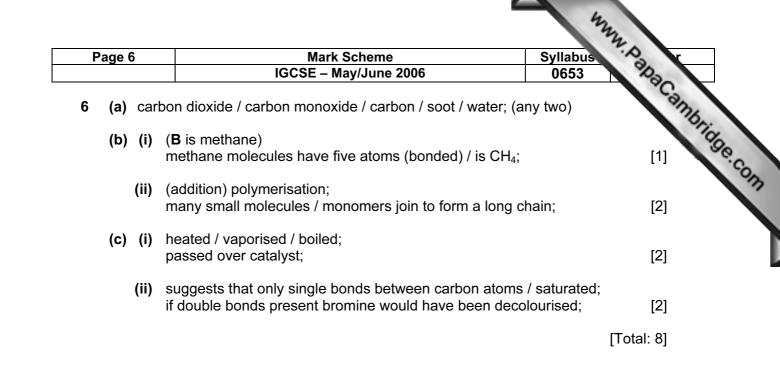
[Total: 6]



Page 3	Mark Scheme S	Syllabus of r
	IGCSE – May/June 2006	0653
3 (a) (Cl)	gas	Syllabus 0653 [1] [2]
(Br)	liquid	1017
(I)	solid;	[1]
(b) (i) 1	our shared pairs;	
	symbols correctly shown for each atom;	[2]
(ii) <i>4</i>	$4\mathbf{C}l_2 + \mathbf{C}\mathbf{H}_4 \rightarrow \mathbf{C}\mathbf{C}l_4 + 4\mathbf{H}\mathbf{C}l;$	[1]
(iii)	(fluorine)	
I	reactivity decreases down Group 7 / owtte;	[1]
(c) (i)	(nucleus of) $Cl - 37$ contains more neutrons than $Cl - 35$;
	2 more;	[2]
(ii)	(12 x 1) + (4 x 35.5);	
	= 154;	[2]

Page 4	Mark Scheme	Syllabus
	IGCSE – May/June 2006	0653
4 (a) (i)	reference to emission;	am
	reference to products;	[2] Orig
(ii)	relatively short half life but not too short;	Syllabus 0653 [2] [1]
(iii)	3 half lives;	
	so 0.2 g;	[2]
(b) (i)	high voltage means low current;	
	this reduces energy losses;	[2]
(ii)	resistance = voltage/current;	
()	= 22 ohms	[2]

Page 5			Mark Scheme Syllab	Syllabus Syllabus	
			IGCSE – May/June 2006	0653	
5	(a)	nucle	us A	- an	
	()	cell w	vall C		On
			oplast none (allow A)		30
		cell si	urface membrane B	Syllabus 0653 Anacan	- 6
			rrect two marks		
			correct one mark		
		one o	or two correct no marks	[2]	
	(b)		ef to water molecules ;		
			vater passes from beaker through ppm ;	 .	
			because more water outside than inside / correct ref to gra		
		5	tarch (molecules) cannot pass through the membrane ;	[3 max]	
		(ii) a	idd iodine (solution) ;		
			orange / brown / yellow ;	[2]	
	(c)	into ro	oot hair ;		
	(-)		s cells in root ;	[2]	



Page 7	Mark Scheme	Syllabus	2
	IGCSE – May/June 2006	0653	Da
	es / plants, stop rain hitting the ground (hard) ; s hold soil in place ;		anacannbridg max]
terrad	ing stops water running down slopes ;	[2	max]
(b) colou	rless / green / small / no petals / dangling anthers / da	ngling stigmas ;	[1]
(c) (i) b	y diffusion ;		
t	nrough wall of small intestine ;		_
r	ef. to villi ;	[2	max]
(ii) p	ancreas;		
S	ecretes insulin ;		
C	auses, cells / liver, to take glucose from the blood ;		[3]
(iii) a	s level moves away from norm ;		
• •	rocess initiated to bring it back ;		
t	ake these points from a specific description		[2]
		[Tota	l· 1∩1
		້າງດາວ	

Page 8		Mark Scheme	Syllabus	2
		IGCSE – May/June 2006	0653	No.
8 (a)	(i)	bonded together; element cannot be simplified and a compound can be be is made from different elements;	contains differe roken into its el [1
	(ii)	Fe ³⁺ ; working refers to charge balance; (reject vague criss cross answers)		[2]
(b)	pre	yer of zinc covers the steel / provides a barrier; vents reaction between steel and oxygen and water; ow correct references to sacrificial protection)		[2]
(c)	(i)	H⁺;		[1]
	(ii)	no more gas evolved;		[1]
	(iii)	grey crystals appear / magnesium reacts and dissolves; a metal displacement reaction occurs / or equation; because magnesium more reactive than zinc;		[3]
			[Tot	tal: 10]

Page 9		ł	Mark Scheme	Syllabus Syllabus
			IGCSE – May/June 2006	0653
9	(a)	(i) :	acceleration;	and
		(ii)	constant speed;	[1] 136
	(b)		under curve = ; (or other suitable) + 25 = 175 m;	Syllabus 0653 [1] [2]
	(c)	equa	al and opposite/ balanced	[1]
	(d)		force = mass x acceleration; = 120 N;	[2]
			power = work/time; = 600 W;	[2]
	(e)		no mark st CoG;	
			e wider than P;	[2]