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UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE O Level

MARK SCHEME for the May/June 2006 question paper

5070 CHEMISTRY

5070/04

Paper 4

maximum raw mark 60

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All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

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Page 2	Mark Scheme	Syllabus	.0
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- **1** A (1)
- 2 (a) (i) To increase the speed of the electrolysis (1)
 - (ii) Oxygen (1) relights a glowing splint (1)
 - (iii) Hydrogen (1) pops in a flame (1) (or vice versa for consequential gas tests)
 - (b) Twice as much gas in Y as X (1)
 - (c) Chlorine (1) bleaches litmus (1) sodium – vigorous reaction, , dissolves, effervescense, gas given off etc. (any 2 – 2 marks) iron – no reaction (1)

[11]

- 3 (a) (i) cream (1)
 - (ii) filtration (1)
 - **(b) (i)** 0.045 (1)
 - (ii) 0.050 (1)
 - (c) $0.045 \times 188 (1) = 8.46 g (1)$
 - (d) $0.050 \times 188 (1) = 9.4 g (1)$

[8]

4 to 8 (b), (b), (b), (d), (b) 1 mark each

[5]

- **9** (a) potassium manganate(VII) cannot oxidise iron(III) or iron(III) cannot be oxidised (1)
 - **(b)** 6.08 g (1)
 - **(c)** pipette (1)
 - (d) (i) green, colourless (1)
 - (ii) purple, pink (1)
 - (e)

26.3	29.4	47.2	[mark row <u>s or columns to</u> the benefit of the
0.0	3.6	21.6	candidate.
26.3	25.8	25.6	One mark for each correct row or column (3)]
Mean value 25	5.7 (1) cm ³		

- **(f)** 0.000514 (1)
- (g) 0.00257 (1)
- **(h)** 0.0257 (1)
- (i) 3.91 g (1)
- (j) 64.25% (1) [14]

Page 3	Mark Scheme	Syllabus
	GCE O Level – May/June 2006	Syllabus 7 7 7 7 7 7 7 7 7 7 7 7 7 7 7 7 7 7 7
10 (a) colou	rless (1) solution	Cambridge
(b) (i) w	hite ppt. (1)	Tab. Ca
(ii) s	oluble in excess (1)	SH SH
(c) (i) w	hite ppt. (1)	

- **(b) (i)** white ppt. (1)
 - (ii) soluble in excess (1)
- (c) (i) white ppt. (1)
 - (ii) soluble in excess (1)
- (d) dilute nitric acid (1)

aqueous lead(II) nitrate or aqueous silver nitrate (1)

yellow ppt. (1)

ZnI₂(1)[9]

- 11 (a) all points plotted correctly (1) smooth curve through all the points (1)
 - **(b)** 2.0 (1)
 - (c) 13.8 (1) indicate extension on graph back to y-axis (1)
 - **(d) (i)** 7.0 (1)
 - (ii) 27.0 (1) cm³
 - (e) (i) reduce volume or evaporate (1) allow to stand (1) filter off the crystals (1)
 - (ii) Molar mass of $Na_2SO_4 = 142 g (1)$ Mass of Na₂SO₄ = 142 x 0.025 x 0.5 = 1.78 g.(1) (for answers (b), (c), and (d) read the candidate's graph)

[12]