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## for the guidance of teachers

## **5070 CHEMISTRY**

5070/32

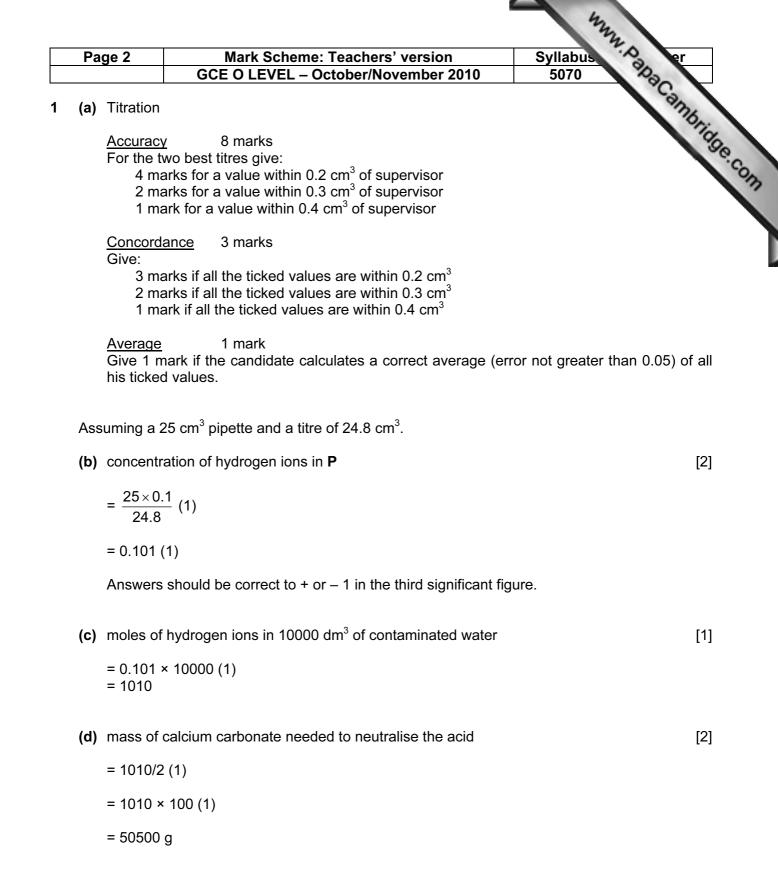
Paper 3 (Practical Test), maximum raw mark 40

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Page 3	Mark Scheme: Teac	Syllabus Syllabus				
	GCE O LEVEL – October/November 2010 5070					
R is aluminium S is potassium nitrate						
Test		Notes				
<b>General po</b> For ppt allow solid,	<b>ints</b> suspension, powder		Syllabus 5070 Sorro Sorr			
	s requires test to be at least pa = bubbles = gas vigorously ev		volved)			
Solutions Colourless r	not equivalent to clear, clear no	t equivalent to colourles	s			
Solution <b>R</b>						
Test 1						
effervescence pops with a lighted splint hydrogen		(1) (1) (1)				
Test 2						
white ppt soluble in excess colourless solution		(1) (1) (1)				
Test 3						
white ppt insoluble in excess		(1) (1)				
Test 4						
pop	ervescence os with a lighted splint drogen	(1) (1) (1)				
	ite ppt uble in excess ourless solution	(1) (1) (1)				
Test 5						
<b>(a)</b> no	reaction	(1)				
blu	/brown solid formed e colour fades ervescence	(1) (1) (1)				

Page 4	Mark Scheme: 1 GCE O LEVEL – Oct			Syllabus Anger 5070
Test 6				Syllabus 5070 n-yellow or colourless reen ppt
(a) liquid turns green		(1)	accept green-yellow or colourless	
(b) green ppt insoluble in excess		(1) (1)	black/dirty green ppt	
Test 7				
turns litmus blue		(1)		
ammonia		(1)		

[20]

[3]

**R** is aluminium/Al (ppt must dissolve in test 2 and ppt must not dissolve in test 3) (1)

**R** is acting as a reducing agent (any green in test **6(a)** or green/black in test **6(b)** (1) **S** contains nitrate or  $NO_3^-$  (test 7 correct – allow alkaline gas, smell of ammonia) (1)

Note: 26 marking points, maximum 23.